**Salty Roads, Take Me Home**

Salted Roads

Today in Science class, Mr. Smith talked about how towns put salt on their roads so that the salt melts the ice. This happens because salt makes it so that liquids freeze at a much lower temperature. Also, towns put sand on their roads so that they are gritty, and gritty ice is not nearly as slippery as ice with nothing on it. He filled a beaker with snow, melted it a little, and then put a bunch of salt in it. Before he did this, we measured the temperature of the snow. It was -1 °C. Then he put the salt in and stirred it, and it melted. We measured the temperature again, and it was -11 °C.

Things I learned:

* H2O is the chemical formula for water.
* The H is hydrogen.
* The 2 is because there are two hydrogen molecules.
* The O is for oxygen. There is only one oxygen molecule.
* The chemical formula for Calcium Chloride CaCl2.
* The chemical formula for Sodium Chloride is NaCl.
* You can melt stuff with microwave energy (hence microwaves).
* Salt rusts things easily.
* This is why cars get rusty, from all the salt in the winter on the roads getting on them.
* In California, it doesn’t get as cold, so they don’t need salt.
* Antifreeze keeps the water inside your car from freezing.

